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**David Baud**

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of the forward reaction is equal to the rate of the  
reverse reaction we learned about rates in unit  
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standard gibbs free energy of 400 kJ/mol while  
the hydrogen bond and van der Waals forces  
have substantially lower values respectively 25  
and 5 kJ/mol given such values which of the  
following is equivalent to the ratio of the  $K_{eq}$   
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web in a chemical equilibrium the forward and reverse reactions occur at equal rates and the concentrations of products and reactants remain constant if we run a reaction in a closed system

so that the products cannot escape we often find the reaction does not give a 100 yield of products

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