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*Census of India, 1891 ,1892*

**Census of India, 1891** E. D. MacLagan,1892

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*Digest of Proceedings ...* Railway Signal Association,1912

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hfus what is the heat of vaporization

formula h<sub>n</sub> hvap what is on the x axis of

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reaction was multiplied by 2 therefore  $\Delta h$

must also be halved 3 answer c explanation

option c is incorrect because 566 kJ mol is

the  $\Delta h$  per mole of o<sub>2</sub> and not for 2 moles

even though options b and d are expressed

in grams they are still correct because 56

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answer key what is the heat change when

472 g of carbon reacts with excess o<sub>2</sub>

according to the following equation how

much heat energy must be expelled by the

body to rid the body of one pound of fat 2s

3o<sub>2</sub> 2so<sub>3</sub> h 7914 kJ is this endothermic or

exothermic 2

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transfer of heat thermal energy in chemical

reactions exothermic transfer of heat from

the system to the surroundings 2h<sub>2</sub>g o<sub>2</sub>g

2h<sub>2</sub>o l energy endothermic the transfer of

heat from the surroundings to the system

energy 2hg s 2hg l o<sub>2</sub>g

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chemical equations that show the

associated enthalpy change enthalpy is an

extensive property depends on amount of

matter present this fact means that the

magnitude of h is directly proportional to

the amount of reactant consumed in the

process

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web q 140 g 4 184 j 16oc g oc q 9372 16 j

35 5 g x 1 mole kcn 0 5453 moles 65 1 g 0

5453 mole moles 5 when a 15 g sample of

srcl<sub>2</sub> dissolves in 80 g of water in a

calorimeter the temperature drops from 42

oc to 15 c calculate the  $\Delta h$  for the process

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explanations 1 a a thermodynamic system

has two means of exchanging energy with

its surroundings heat flow and pressure

volume work heat flow q is the transfer of

thermal energy due to temperature difference work  $w$  is macroscopic energy transfer between the system and its surroundings that occurs

### **thermochemistry practice problems**

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web in this set of practice questions we will summarize the main concepts of thermochemistry such as the relationship between internal energy work and heat exothermic and endothermic process heat capacity constant pressure calorimetry constant volume calorimetry the enthalpy the standard enthalpies read more

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web jan 10 2021  $\Delta H = 27.4 \text{ kJ mol}^{-1}$  and  $\Delta S = 29.8 \text{ J K}^{-1}$  sulfur trioxide  $\text{SO}_3$  is produced in large quantities in the industrial synthesis

of sulfuric acid sulfur dioxide is converted to sulfur trioxide by reaction with oxygen gas write a balanced chemical equation for the reaction of  $\text{SO}_2$  with  $\text{O}_2$  and determine its  $\Delta G$

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